Jl
ill

heat .of the 0 (as above found, i.e.y 28 calories) + the heat resulting from the formation of xGO^ consequently from formula (20), 4,344# ; minus the heat which would be associated with the production of x cubic metre H, which from (21) amounts to — 2,585 y. Therefore Q=28+4,344a;--2,585;!/=279 calories.

It admits of being approached further from the formula for steam-air gas, from which C02=ff, G0=zero, H=y,

«=S-762a?— I,88l2/; and from

wherefrom

1— ^—7/=3'762^— l,881y, .'. 0-881j/=4-762a;— 1,

;— 1-135.

Placing these values in the equation for Q one obtains

279=Q— 28+4,344^— 2,585(5-405^— 1135),

.". 279=2,962—9,628^,

whence #=0*279 cubic metre COa

from ^=5-405ar— 1-185, 7/=0-373 „ H

from z=^l—x—y ^==0-348 _ „ N

1*000 cubic metre gas.

The combustible value of this gas calculates out thus : — 0-373 cubic metre Hx 2,620=977 calories.

It results herefrom that the combustible value of the gas is extraordinarily low, even being inferior to the combustible value of pure air-gas. Thus on this ground steam-air gas with COa formation appears as of minor value.

Nevertheless, possibly the utilisation of the fuel could be achieved better, and in order to investigate this, the amounts